Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_\_\_\_\_

**CLASSWORK: Mole to Atom/Particles/Molecule Conversion**

For the following problems 1) Circle the given AND underline the unknown 2) decide whether to convert **moles 🡪 particles/atoms/molecules ,**  **OR particles/atoms/molecules 🡪 moles** 3) set up your T-Chart 4) solve for what is requested making sure to include units. **You must show all work, and every item must have a unit!** (5 pts each)

1. Calculate how many atoms of calcium chloride (CaCl2) are present in a 0.25 mole sample.
2. Calculate how many moles of sulfur are present 7.85 x 1024 atoms of sulfur.
3. How many moles of copper (II) chlorate (Cu(ClO3)2) are present in 5.50 x 1023 atoms of copper (II) chlorate?
4. How many molecules of water (H2O) are present in 3.50 moles of water?

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_\_\_\_\_

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