

# Notes: Periodic Table Basics

Each element on the Periodic Table (PT) is different based on the number of protons in the nucleus. (Protons are the atoms DNA)

6	← atomic number
C	← atomic symbol
12.011	← atomic mass
Carbon	← element name

- The atomic number is the number of protons in an element's nucleus.

↳ The number of protons NEVER changes because it's like the atoms DNA!!

- The atomic symbol is a short version of the element name.

- The atomic mass is the mass of an atom

↳ Atomic mass = # protons + # neutrons

# neutrons = atomic mass - # protons

↳ The atomic mass on the periodic table is NOT a whole number because it is an average of ALL the known atoms of that element

↳ each atom is known as an ISOTOPE

- A neutral atom has the same # of protons & electrons

Ex:

13
Al
26.982
Aluminum

Atomic #: 13

Atomic mass: 26.982 ≈ 27

#P: 13; #E: 13; #N: 27 - 13 = 14