Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_\_\_\_\_

**CLASSWORK: Mole to Grams Conversion**

**Write your observations for the two beakers at the front of the room.**

Salt

Sugar

**What is the difference that you can see between the two?**

**Why do you think that is?**

**Define:**

Molar mass: The of mole of a certain element or compound

The values for molar mass can be found

Units for molar mass are:

**Now that you have completed some whiteboard practice complete some practice molar mass problems**

1. NO2 2) CO2

3) CaCl2 4) C6H12O6

Back to the demo: Draw the T- Chart

2. How many grams are in 5.69 moles of uranium?

3. How many moles are in 8.9 grams of neon?

4. How many grams are in 421 grams of sodium?

**Practice: use the space under this sheet for more room**

5. How many grams are in 5.8 moles of H2O

6. How many grams are in 0.75 moles of CaCl2

7. How many moles are in 34 grams of N