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| **Date** | **Topic** | **Homework** |
| ***Pre-Accelerated*** | ***Honors*** |
| M – Feb 1 | Intro to Covalent Bonding | Questions 1-8***Study Polyatomics!*** | Questions 1-8***Study Polyatomics!*** |
| T – Feb 2 | Compounds in Covalent Bonds**Polyatomics Re-Test!** | Write lab procedure | Write lab procedure |
| W – Feb 3 | Ionic v. Covalent Compounds Lab | Lab Conclusion Questions | Lab Conclusion Questions |
| Th – Feb 4 | Naming Covalent Compounds | Questions 9-10 (a-f) | Questions 9-10 (c-h) |
| F – Feb 5 | Lewis Dot Structures | Question 11 | Question 11 |

**\*\*\*Complete the following problems on a separate sheet of paper in your chemistry notebook.**

1. What is the difference between a compound and a molecule?

2. Why can’t two nonmetals form an ionic bond?

3. Define a covalent bond in your own words.

4. Determine if the following compounds are ionic, covalent, or neither. If they are ionic, name them.

a. H2O

 b. MgO

 c. NaNO3

 d. NO31-

 e. FeCl2

 f. NO2

 g. SiCl4

 h. Ni2(SO4)3

5. How many electrons are shared in a single covalent bond? A double covalent bond? A triple covalent bond?

6. In covalent bonding, how many **pairs** of electrons are shared in a single bond? A double bond? A triple bond?

7. What are the four properties of covalent compounds?

8. How are ionic and covalent compounds different?

9. Name the following compounds:

* 1. N2O5 c. CH4 e. BF3 g. CO
	2. SeBr6 d. C6H6 f. NO2 h. BrO3
1. Write the formulas for the following compounds:
	1. dihydrogen monoxide
	2. phosphorous trihydride
	3. carbon tetrachloride
	4. carbon dioxide
	5. dichlorine heptaoxide
	6. disilicon hexabromide
	7. nitrogen dioxide
	8. iodine heptafluoride

11. Fill in the following table: (continued on the back)

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| **A.) Electrons** | **Draw the Lewis Dot for: H2S** | **B.) Electrons** | **Draw the Lewis Dot for: CH4** |
|  |  |  |  |
| **C.) Electrons** | **Draw the Lewis Dot for: N2** | **D.) Electrons** | **Draw the Lewis Dot for: CO2** |
|  |  |  |  |
| **E.) Electrons** | **Draw the Lewis Dot for: CO** | **F.) Electrons** | **Draw the Lewis Dot for: OF2** |
|  |  |  |  |
| **G.) Electrons** | **Draw the Lewis Dot for: N2O** | **H.) Electrons** | **Draw the Lewis Dot for: O3** |
|  |  |  |  |
| **I.) Electrons** | **Draw the Lewis Dot for: NO31-** | **J.) Electrons** | **Draw the Lewis Dot for: NO21-** |
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