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| **Date** | **Topic** | **Homework** |
| **CB** | **PA** |
| M – Feb 2 | NO SCHOOL – Snow Day | NA | NA |
| T – Feb 3 | Introduction to Covalent Bonding | 1-2, 4a-f, 5-6 | 1-6 (all) |
| W – Feb 4 | Covalent Bonds & Diatomic Elements | Lab Procedure Write Up |
| Th – Feb 5 | Comparing Ionic & Covalent Bonds Lab | Lab Write Up (Conclusion Questions) |
| F – Feb 6 | Naming Covalent Compounds | 7, 8a-f, 9a-f | 7-9 (all) |

**\*To receive credit, complete the following problems on a separate sheet of paper behind your agenda and the catalyst in your chemistry notebook.**

1. What is a covalent bond? What is happening with electrons in a covalent bond?
2. Review. What is an ionic bond? What is happening with electrons in an ionic bond?
3. Why won’t helium or neon normally bond with other atoms?
4. Determine if the following pairs of atoms will form an ionic bond, a covalent bond, or a metallic bond.
	1. Na and Br
	2. Cl and O
	3. H and Cl
	4. Fe and C
	5. Cu and Zn
	6. C and O
	7. Co and O
	8. Li and K
	9. H and F
	10. Ag and Au
5. How many electrons are shared in a single covalent bond? A double covalent bond? A triple covalent bond?
6. Rank the following from strongest to weakest: single covalent bond, double covalent bond, ionic bond, triple covalent bond.
7. Draw this chart onto your notebook paper with your answers to the agenda questions and fill it out. List AT LEAST 3 PROPERTIES EACH.

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| **Properties of Ionic Compounds** | **Properties of Covalent Compounds** |
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1. Write the formulas for the following compounds:
	1. dihydrogen monoxide
	2. phosphorous trihydride
	3. carbon tetrachloride
	4. carbon dioxide
	5. dichlorine heptaoxide
	6. disilicon hexabromide
	7. nitrogen trioxide
	8. triphosphorus pentoxide
2. Name the following compounds:
	1. N2O5
	2. CH4
	3. NO2
	4. C6H6
	5. SeBr6
	6. BF3
	7. CO
	8. BrO3