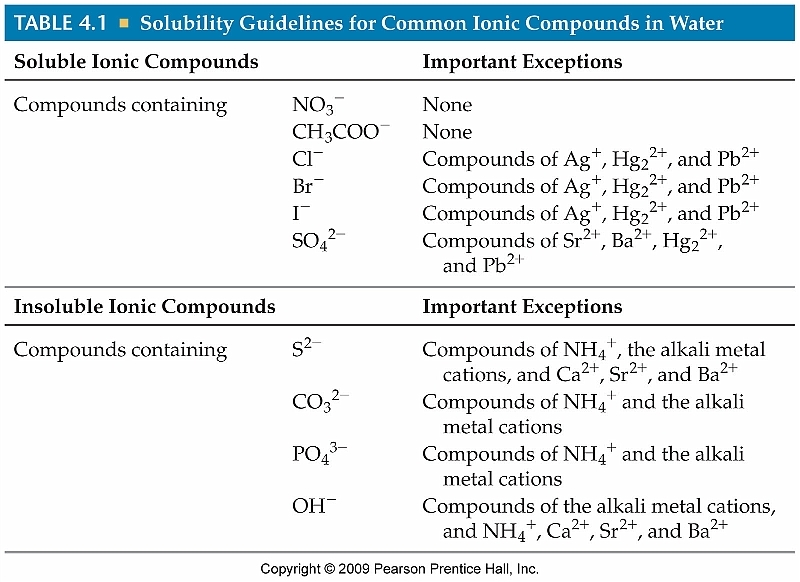
Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_\_\_

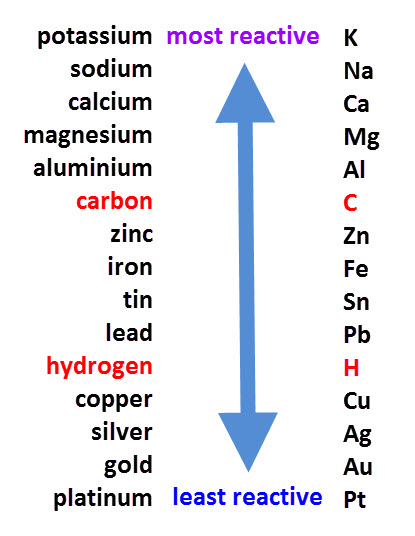
**WEEK 27 AGENDA: Unit 6 website: kachemistry.weebly.com**

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| --- | --- | --- |
| **Date** | **Topic** | **Homework** |
| **Pre-Accelerated** |
| M – Mar 16 | Solubility Rules Lab | Post-Lab #1, #1-6 |
| T – Mar 17 | Solubility Rules Cont. | Completed Solubility Lab Packet due tomorrow |
| W – Mar 18 | QUIZ, PLAN Review | 7-14 |
| Th – Mar 19 | Reactivity Series Lab | 15-19 |
| F – Mar 20 | Reactivity Series Cont. | 20-25 |

**\*Complete the following problems on a separate sheet of paper in your chemistry notebook.**

|  |  |
| --- | --- |
| 1.) Type of Chemical Rxn: | 2.) Type of Chemical Rxn: |
| \_\_\_\_\_\_C10H16(l)+ \_\_\_\_O2(g)🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? | \_\_\_\_\_Ca3(PO4)2(aq) + \_\_\_\_\_SiO2(aq) 🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? |
| 3.) Type of Chemical Rxn: | 4.) Type of Chemical Rxn: |
| \_\_\_\_\_\_(NH4)2CO3(aq) + \_\_\_\_\_ HCl 🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? | \_\_\_\_\_Al(s) + \_\_\_\_HCl🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? |
| 5.) Type of Chemical Rxn: | 6.) Type of Chemical Rxn: |
| \_\_\_LiClO*(aq)* + \_\_\_KOH*(aq)* 🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? | \_\_\_CuCl 🡪  ☐Compounds: Charges Equal  ☐Elements: Diatomics? |

Write the balanced equation for the following reactions AND determine the state of matter for each product: (Use the table at right)

1. Zn + CuSO4 🡪 Cu + ZnSO4
2. Mg + AgNO3 🡪 Ag + Mg(NO3)2
3. Cl2 + NaBr🡪 NaCl + Br2
4. CaO + H2O 🡪 Ca(OH)2
5. Ca(ClO3)2 🡪 CaCl2 + O2
6. AgNO3 + K2SO4 🡪 Ag2SO4 + KNO3
7. Zn*(s)* + CuBr2*(aq)* 🡪 ZnBr2 + Cu
8. C8H12+ O2 🡪 CO2 + H2O
9. Determine if each of the follow substances would be soluble (aq) or insoluble (s) in solution:
   1. KBr
   2. PbCO3
   3. BSO3
   4. zinc hydroxide
   5. sodium acetate
   6. silver iodide
   7. cadmium (II) sulfide
   8. zinc carbonate
10. Name three things that always stay the same in a chemical reaction.
11. What changes in a chemical reaction?
12. Define “reactant” and “product” in your own words.
13. How many atoms of each type of element are present in the following compounds:
    1. BeSO4 e. Al(ClO2)3
    2. 3BeSO4 f. 5Ba3(PO4)2
    3. KClO4 g. Li2MnO4
    4. 7KClO4 h. 6Li2MnO4
14. Balance the following: (SHOW ALL WORK!!)
    1. \_\_\_CaCO3 + \_\_\_H3PO4 🡪 \_\_\_Ca3(PO4)2 + \_\_\_H2CO3
    2. \_\_\_NaClO3 🡪 \_\_\_NaCl + \_\_\_O2
    3. \_\_\_KOH + \_\_\_Co3(PO4)2 🡪 \_\_\_K3PO4 + \_\_\_Co(OH)2
15. Based on the reactivity series diagram to the right, circle the more reactive metal:
    1. Platinum or Zinc
    2. Carbon or Hydrogen
    3. Magnesium or Gold
    4. Silver or Gold
    5. Calcium or Zinc
16. Explain why sodium can replace calcium in a compound but cannot replace potassium.
17. Explain, in terms of reactivity, why most jewelry is made of silver, gold or platinum.
18. Will the following reactions take place? (If Yes, write the new products and balance the equation)
    1. Zn + Pb(NO3)2
    2. Fe + ZnSO4
    3. Pb + Cu(NO3)2
    4. Au + Pb(NO3)2
    5. Mg + Ca(NO3)2
    6. Ca + H2O